**Chemistry 20 – Lesson 5**

**The mole**

**/58**

**Part A**

1.



/2

2.



/2

3.



/2

4.



/2

5.



/2

6.



/2

7.

/2



8.



/2

9.



/2

10.



/2

11.



/2

12.



/2

**Part B**

****1.



/3



****

2.

/3



****3.



/3



****4.



/3



****5.



/3



****6.



/3



****7.



/3



8.

****

/3



****9.

****

/7





 ∴70 g of helium contains a greater number of moles than 250 g of water

10. If we calculate the molar mass of the gas we can use the periodic table to identify it.

/3





 The inert gas with a molar mass of 84 g/mol is **krypton**.